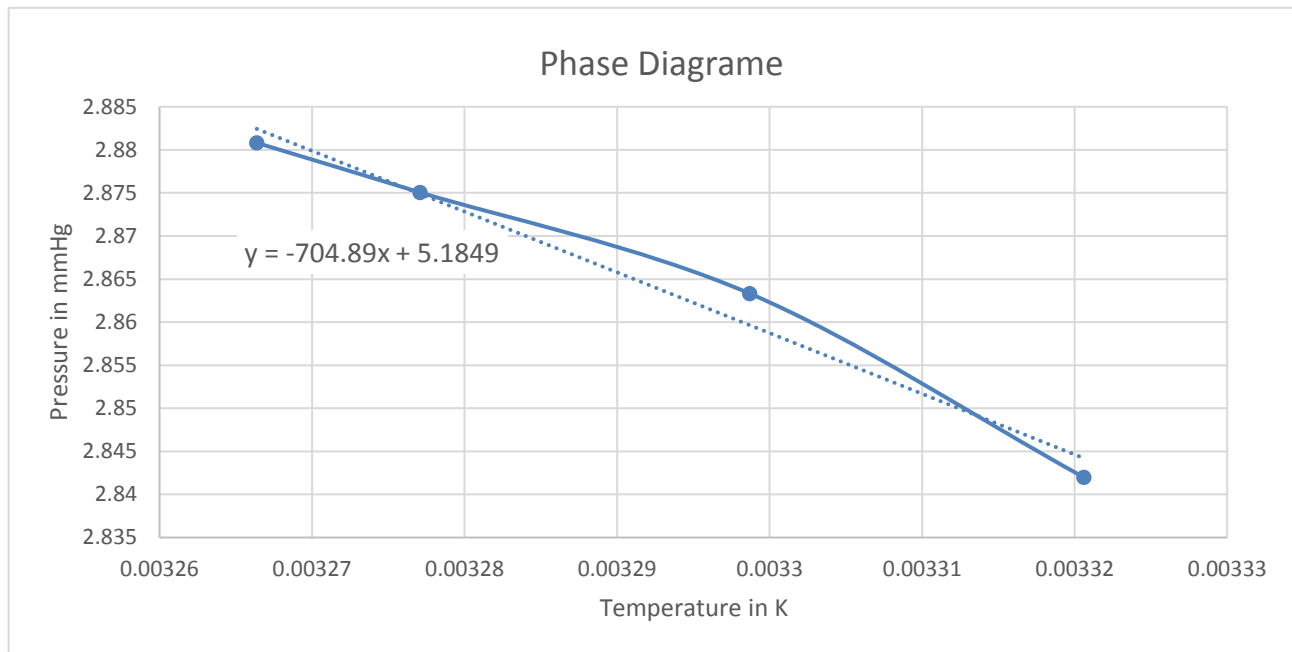


Q1) By using Clausius–Clapeyron relation $\text{Log}(P) = \frac{-\Delta H_{vap}}{2.303 * R * T} + C$

P (mmHg)	Log(P (mmHg))	T (°C)	T (K)	1/T (K)
760	2.880813592	33	306.15	0.003266
750	2.875061263	32	305.15	0.003277
730	2.86332286	30	303.15	0.003299
695	2.841984805	28	301.15	0.003321

By representing the data and taking the best fit :



	R=	1.987	Cal/(gmol*K)
part 2	C =	intercept =	5.1849
	(-ΔHvap)/(2.303*R) =	Slope (K) =	-704.89
	(-ΔHvap) =	Slope*R*2.303	-3225.62
part 1	ΔHvap =	3225.62	Cal/(gmol)